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where  $T_0$  is the

freezing point of the  
solvent. If BC in Fig.

1st-1 is taken to be a

straight line, the slope

$dp_A/dT$  may be

replaced by  $CB/BD$ ,

i.e., by  $(p_A - p_s)/\Delta T$ ,

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so that  $p_p T H_p R T$

$\Delta_{\text{smvA}} 0.2 \Delta = \Delta \Delta 0$

Eq. 1st.4 The slope of the vapor pressure

curve of the solid

solvent is also given

by another form of the

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The dilution formula  
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determine the molarity of a solution after it has been diluted. The formula is:  $M_1 V_1 = M_2 V_2$  Where  $M_1$  and  $V_1$  are the molarity and volume before a dilution, and  $M_2$  and  $V_2$  are the molarity and volume after a dilution. Consider this example from the 2006 NECTA:

"Hydrochloric acid

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solution made by  
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1st2 Task Freezing

and melting point will

be determined by

using the apparatus

described on Fig 1st-2

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Packing models of (a)  
side-by-side layer, (b)

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hexagonal layer, (c)  
simple cubic layer, (d)  
body centered cubic.  
Determine the

maximum amount of  
balls that can touch  
another one ball  
(center- ball) in the  
same layer for each  
arrangement.

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