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will discover  
that many  
reactions and  
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Châtelier's principle to explain how various factors affect chemical equilibria. You will calculate equilibrium concentrations of reactants and products using the equilibrium constant

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 $\text{Mg}(\text{OH})_2$ , is a  
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Concentration,  
Ion Products,  
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1. In a bottle of unopened cola, the  $\text{CO}_2$  gas dissolved in the liquid is in equilibrium with the  $\text{CO}_2$  gas above the

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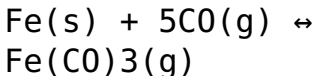
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$\frac{[\text{Fe}(\text{CO})_5]}{[\text{Fe}(\text{CO})_3]}$  /  
[CO]<sup>5</sup>. If the  
equation  
 $\text{CH}_3\text{OH}(\text{g}) + 101\text{kJ}$   
 $\rightleftharpoons \text{CO}(\text{g}) +$   
 $2\text{H}_2(\text{g})$  is for a  
system at  
equilibrium,  
increasing the  
temperature will  
cause. [CH<sub>3</sub>OH]  
to decrease and  
[CO] and [H<sub>2</sub>] to  
increase.



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A solution  
equilibrium  
occurs when a  
solid substance  
is in a  
saturated  
solution. At  
this point, the  
rate of  
dissolution is

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equal to the rate of recrystallization.

Although these are all different types of transformations, most of the rules regarding equilibrium apply to any situation in which a process

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occurs  
reversibly.

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reaction in

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products can

react to reform

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reaction equals the rate of its reverse reaction and the concentrations of its products and reactants remain unchanged.

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1.  $A + B \rightarrow C + D$   
(forward  
reaction)  $C + D$   
 $\rightarrow A + B$  (reverse  
reaction)

Equilibrium

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(forward rate = reverse rate) remain constant. The ratio of the mathematical product  $[C]^x \times [D]^y$  to the mathematical product  $[A]^n \times [B]^m$  for this reaction has a definite value at a given temperature.

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“Equilibrium” is

the seventh

chapter in the

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textbook.

Several important concepts such as equilibrium constants, buffer solutions, and the common-ion effect is explained in this chapter.

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Similarly, in Chapter 13, we discussed saturated solutions, another example of a physical equilibrium, in which the rate of dissolution

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of a solute is the same as the rate at which it crystallizes from solution.

In this chapter, we describe the methods chemists use to quantitatively describe the composition of chemical systems at equilibrium

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Reaction Rates  
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In layman's  
terms,

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equilibrium is defined as a state of balance due to equal reactions of opposing forces, and today we'll be talking all about it with regards to the scientific study of chemistry, focusing on such topics as

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Chemical reaction rates.

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In this chapter,  
we will learn  
about the types  
of equilibrium,  
characteristics  
of chemical  
equilibrium,  
reversible and

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irreversible  
reactions, pH  
scale, the study  
of pH and pOH  
and much more.

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