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13. *Molecular*

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~~Sp³, Organic~~

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*Kernberg's main
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A bond in which
the bonding
electrons are
most likely to
be found in the

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sausage-shaped regions above and below the nuclei of the bonded atoms. (sigma bond, pi bond, VSEPR theory, hybridization, linear molecule)

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Answers (sigma

bond, pi bond,

VSEPR theory,

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sigma bond. A

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be found in the
sausage-shaped
regions above
and below the
nuclei of the
bonded atoms.

(sigma bond, pi
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pi bond.

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9 Theories of
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9 3 9 3 A

covalent bond is
the result of

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the overlap of
orbitals on
adjacent atoms

The bonding
region is the
location between
the atomic
nuclei where
electrons occupy
the overlapping

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8.2 Part A

Completion 1.

stable electron

covalent shared

single unshared

pairs

double/triple

coordinate

covalent bond

Energy bond

dissociation

energy

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8.2 Part A

Completion 1.

stable electron

...

just as an

atomic orbital

belongs to an

atom, a

molecular

orbital belongs

to a molecule.

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bonding orbital.
a molecular
orbital that can
be occupied by
two electrons of
a covalent bond.
sigma bond. two
atomic orbitals
combine to form
a molecular
orbital that is
symmetrical
around the axis
connecting two

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atomic nuclei.

pi bond.

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chemistry 8.3

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9-3 9-3 A

covalent bond is

the result of

the overlap of

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orbitals on adjacent atoms. The bonding region is the location between the atomic nuclei, where electrons occupy the overlapping orbitals. For example, consider the covalent bond in hydrogen,

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H₂ (Figure 9.2) .

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Objectives

Identify the
difference

between atomic
and molecular
orbitals ‘

Describe
how VSEPR theory
helps predict
the shapes of
molecules

Identify the
ways in which
orbital

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Hybridization is
useful in
describing
molecules

Vocabulary ' molecular
orbitals • pi
bond VSEPR
theory

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An interatomic distance is reached at which the potential energy is at a minimum. If the distance decreases beyond this point, the potential energy increases because of increased repulsion. Thus,

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a chemical bond forms with a bond length equal to the interatomic distance at which the potential energy is a minimum.

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